

**CHEMISTRY**

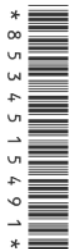
**9701/11**

Paper 1 Multiple Choice

**October/November 2018**

**1 hour**

Additional Materials: Multiple Choice Answer Sheet  
Soft clean eraser  
Soft pencil (type B or HB is recommended)  
Data Booklet



**READ THESE INSTRUCTIONS FIRST**

Write in soft pencil.

Do not use staples, paper clips, glue or correction fluid.

Write your name, Centre number and candidate number on the Answer Sheet in the spaces provided unless this has been done for you.

**DO NOT WRITE IN ANY BARCODES.**

There are **forty** questions on this paper. Answer **all** questions. For each question there are four possible answers **A, B, C** and **D**.

Choose the **one** you consider correct and record your choice in **soft pencil** on the separate Answer Sheet.

**Read the instructions on the Answer Sheet very carefully.**

Each correct answer will score one mark. A mark will not be deducted for a wrong answer.

Any rough working should be done in this booklet.

Electronic calculators may be used.

This document consists of **16** printed pages.

## Section A

For each question there are four possible answers, **A**, **B**, **C** and **D**. Choose the **one** you consider to be correct.

Use of the Data Booklet may be appropriate for some questions.

- 1 The first four ionisation energies for element X are shown in the table.

ionisation energy	1st	2nd	3rd	4th
value / $\text{kJ mol}^{-1}$	577	1980	2960	6190

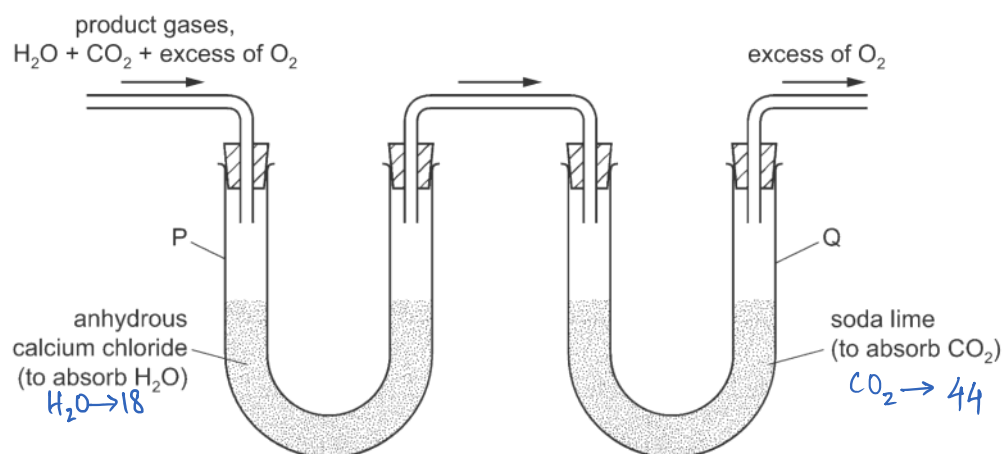
Which ion of X is produced by removing an electron from a filled shell?

- A  $X^+$       B  $X^{2+}$       C  $X^{3+}$       **D**  $X^{4+}$

- 2 What is a basic assumption of the kinetic theory, as applied to an ideal gas?

- A** Collisions between gas molecules are elastic.  
~~B~~ Each gas molecule occupies a finite volume.  
~~C~~ Gases consist of particles that experience the force of gravity.  
~~D~~ Gas molecules attract each other with weak intermolecular forces.

- 3 A sample of the hydrocarbon  $\text{C}_6\text{H}_{12}$  is completely burned in dry oxygen and the product gases are collected as shown.



The increases in mass of the collecting vessels P and Q are  $M_P$  and  $M_Q$ , respectively.

What is the ratio  $M_P / M_Q$ ?

- A**  $0.41 \frac{18}{44}$       B 0.82      C 1.2      D 2.4

- 4 5.0g samples of the carbonates of barium, copper, lithium and magnesium are decomposed to the metal oxides and carbon dioxide.

For which compound is there the greatest loss in mass?

- A barium carbonate  $\text{BaCO}_3 \rightarrow \text{BaO} + \text{CO}_2$   $5\text{g} \rightarrow 3.885\text{g}$   
 B copper(II) carbonate  $\text{CuCO}_3 \rightarrow \text{CuO} + \text{CO}_2$   $5\text{g} \rightarrow 3.2186\text{g}$   
 C lithium carbonate  $\text{Li}_2\text{CO}_3 \rightarrow \text{Li}_2\text{O} + \text{CO}_2$   $5\text{g} \rightarrow 2.019\text{g}$   
 D magnesium carbonate  $\text{MgCO}_3 \rightarrow \text{MgO} + \text{CO}_2$   $5\text{g} \rightarrow 2.3903\text{g}$

- 5 In this question you should assume methane behaves as an ideal gas.

The gas laws can be summarised in the ideal gas equation below.

$$pV = nRT \quad 1.03 \times 10^5 \times 5.37 \times 10^{-3} = \frac{\text{mass}}{16} \times 8.314 \times (333)$$

The volume of a sample of methane is measured at a temperature of  $60^\circ\text{C}$  and a pressure of  $103\text{ kPa}$ . The volume measured is  $5.37 \times 10^{-3}\text{ m}^3$ .

What is the mass of the sample of methane, given to two significant figures?

- A 0.0032g    B 0.018g    C 3.2g    D 18g
- 6 A butane burner is used to heat water. The  $M_r$  of butane is 58.

- $\Delta H_c^\ominus$  of butane is  $-2877\text{ kJ mol}^{-1}$ .
- 250g of water is heated from  $12^\circ\text{C}$  to  $100^\circ\text{C}$ .
- The burner transfers 47% of the heat released from the burning fuel to the water.

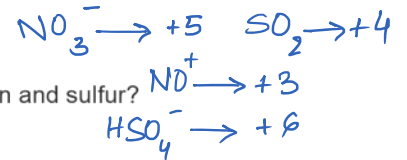
Assume that the butane undergoes complete combustion and none of the water evaporates.

What is the minimum mass of butane that must be burnt?

- A 0.068g    B 1.85g    C 3.94g    D 4.48g

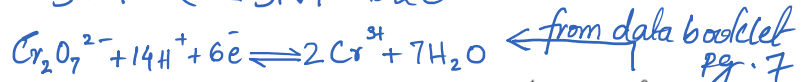
Complete combustion of 1 mole of butane,  $-2877\text{ kJ}$  of energy releases  
 To raise the temp. of 250g of water from  $12^\circ\text{C}$  to  $100^\circ\text{C}$ , energy required  
 is  $250 \times 4.18 \times (100 - 12) = 91960\text{ J}$   
 But this energy was just 47% of the total heat produced by burner  
 Total heat by burner is  $\frac{91960 \times 100}{47} = 1.9566 \times 10^5$   
 $1\text{ mole of butane} \rightarrow 2877 \times 1000\text{ J}$   
 $0.068\text{ moles} \rightarrow 1.9566 \times 10^5$   
 $0.068 \times 58 = 3.944\text{ g of butane}$

- 7 Nitric acid is known to take part in the oxidation of atmospheric sulfur dioxide. One possible reaction is shown.



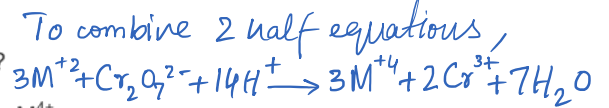
Which row shows the correct changes in oxidation numbers of nitrogen and sulfur?

	nitrogen	sulfur
A	-3	+3
<b>B</b>	-2	+2
C	-2	+3
D	-1	+2



- 8 A transition metal ion,  $\text{M}^{2+}$ , reacts with acidified dichromate(VI) ions to form  $\text{M}^{4+}$  ions,  $\text{Cr}^{3+}$  ions and  $\text{H}_2\text{O}$ .

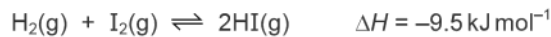
Which equation correctly represents this reaction?



- A  $\text{Cr}_2\text{O}_7^{2-} + 14\text{H}^+ + \text{M}^{2+} \rightarrow 2\text{Cr}^{3+} + 7\text{H}_2\text{O} + \text{M}^{4+}$
- B  $\text{Cr}_2\text{O}_7^{2-} + 14\text{H}^+ + 2\text{M}^{2+} \rightarrow 2\text{Cr}^{3+} + 7\text{H}_2\text{O} + 2\text{M}^{4+}$
- C**  $\text{Cr}_2\text{O}_7^{2-} + 14\text{H}^+ + 3\text{M}^{2+} \rightarrow 2\text{Cr}^{3+} + 7\text{H}_2\text{O} + 3\text{M}^{4+}$
- D  $\text{Cr}_2\text{O}_7^{2-} + 14\text{H}^+ + 6\text{M}^{2+} \rightarrow 2\text{Cr}^{3+} + 7\text{H}_2\text{O} + 6\text{M}^{4+}$

- 9 In this question you should assume that all gases behave ideally.

Hydrogen and iodine react reversibly in the following reaction. The system reaches dynamic equilibrium.

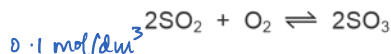


Which statement **must** be true for the  $K_p$  of this equilibrium to be constant?

- A The partial pressures of  $\text{H}_2$ ,  $\text{I}_2$  and  $\text{HI}$  are equal.
- B The external pressure is constant.
- C The forward and reverse reactions have stopped.
- D** The temperature is constant.

Equilibrium constants are only changed when temperature changes

- 10  $0.2 \text{ mol/dm}^3$   $0.2 \text{ mol/dm}^3$   
 0.200 mol of sulfur dioxide and 0.200 mol of oxygen are placed in a 1.00 dm<sup>3</sup> sealed container. The gases are allowed to react until equilibrium is reached.



At equilibrium there is  $0.100 \text{ mol}$  of  $\text{SO}_3$  in the container.

What is the value of  $K_c$ ?

- A  $0.150 \text{ mol dm}^{-3}$   
 B  $0.800 \text{ mol dm}^{-3}$   
 C  $1.25 \text{ mol}^{-1} \text{ dm}^3$   
 D  $6.67 \text{ mol}^{-1} \text{ dm}^3$

	$2\text{SO}_2$	$\text{O}_2$	$2\text{SO}_3$
initial	0.2	0.2	0
change	-0.1	-0.05	+0.1
At equilibrium	0.1	0.15	0.1

$$K_c = \frac{[0.1]^2}{[0.1]^2 [0.15]} = 6.67 \text{ mol}^{-1} \text{ dm}^3$$

There is increase in moles of  $\text{SO}_3$  by 0.1 so  $\text{SO}_2$  should decrease by same amount since they've equal mole ratio but  $\text{O}_2$  will decrease by 0.1 bc2 of mole ratio of 1:2

- 11 Two reactions are shown.



In reaction 1, a finely powdered iron catalyst is used.

In reaction 2, a vaporised  $\text{CCl}_4$  tetrachloromethane catalyst in ultraviolet light is used.

Which statement about the catalysts used is correct?

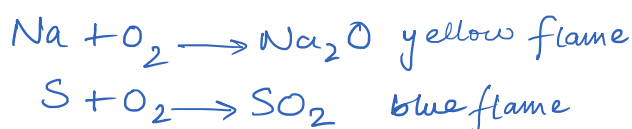
- A Both reaction 1 and reaction 2 use a heterogeneous catalyst.  
 B Both reaction 1 and reaction 2 use a homogeneous catalyst.  
 C Reaction 1 uses a heterogeneous catalyst and reaction 2 uses a homogeneous catalyst.  
 D Reaction 1 uses a homogeneous catalyst and reaction 2 uses a heterogeneous catalyst.

- 12 Sodium and sulfur are burned separately in oxygen.

Each reaction has a distinctive coloured flame.

Which row is correct?

	Na + O <sub>2</sub>	S + O <sub>2</sub>
<del>A</del>	white	blue ✓
B	white	yellow
<input checked="" type="radio"/> C	yellow	blue ✓
D	yellow	yellow



13 X and Y are elements in Period 3 of the Periodic Table.

- The oxide of X is a solid at room temperature. This oxide has a giant structure.
- The chloride of X does not react with water.
- Argon is the only element in Period 3 with a lower melting point than Y.

ionic lattice  
 $\text{Na}_2\text{O}, \text{MgO}, \text{Al}_2\text{O}_3$  or  $\text{SiO}_2$   
 giant molecular

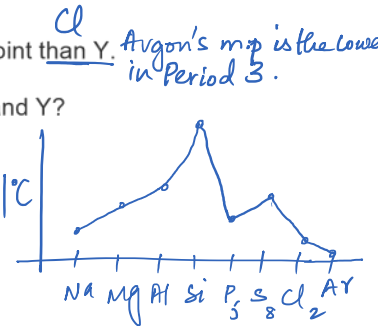
What could be the formula of a compound formed between elements X and Y?



14 Which row describes the structure and bonding of  $\text{SiO}_2$  and  $\text{SiCl}_4$ ?

	$\text{SiO}_2$		$\text{SiCl}_4$	
	bonding	structure	bonding	structure
A	covalent	giant	covalent	giant
<b>B</b>	covalent	giant	covalent	simple
C	ionic	giant	covalent	giant
D	ionic	giant	covalent	simple

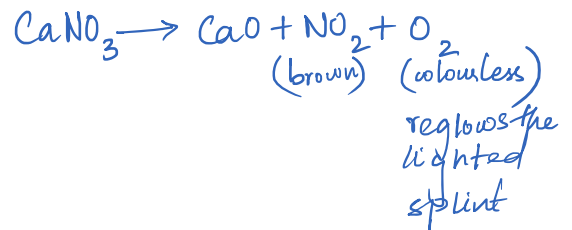
$\text{SiO}_2 \rightarrow$  giant molecular (covalent)  
 $\text{SiCl}_4 \rightarrow$  simple molecular (covalent)



15 A sample of anhydrous calcium nitrate is placed in a test-tube and heated in a roaring Bunsen flame until it decomposes. The description of the gas in the test-tube is then noted. A glowing splint is then put into the test-tube and any changes are noted.

Which observations are correct?

	description of the gas in the test-tube	result of glowing splint test
A	brown	the splint goes out
<b>B</b>	brown	the splint relights
C	colourless	the splint goes out
D	colourless	the splint relights



- 16 Which row correctly describes the properties of the halogens as Group 17 is descended from chlorine to iodine?

	volatility	strength as oxidising agent
<b>A</b>	decreases	decreases
B	decreases	increases
C	increases	decreases
D	increases	increases

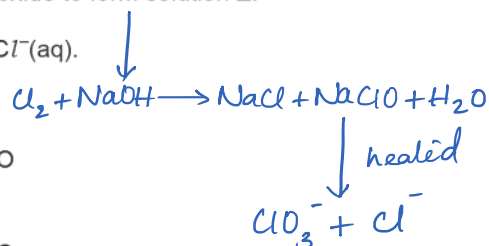
Oxidising power ↓ as we descend the group & volatility decreases too.  
Reducing power increases as we descend group 17

- 17 Reaction 1: chlorine reacts with cold aqueous sodium hydroxide to form solution Z.

Reaction 2: solution Z is heated and forms  $\text{ClO}_3^-$ (aq) and  $\text{Cl}^-$ (aq).

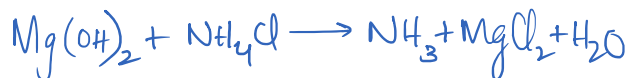
Which equations represent reaction 1 and reaction 2?

- ~~A~~ reaction 1  $2\text{Cl}_2 + 4\text{OH}^- \rightarrow \text{ClO}_2^- + 3\text{Cl}^- + 2\text{H}_2\text{O}$   
 reaction 2  $3\text{ClO}_2^- \rightarrow 2\text{ClO}_3^- + \text{Cl}^-$
- ~~B~~ reaction 1  $2\text{Cl}_2 + 4\text{OH}^- \rightarrow \text{ClO}_2^- + 3\text{Cl}^- + 2\text{H}_2\text{O}$   
 reaction 2  $3\text{ClO}^- \rightarrow \text{ClO}_3^- + 2\text{Cl}^-$
- ~~C~~ reaction 1  $\text{Cl}_2 + 2\text{OH}^- \rightarrow \text{ClO}^- + \text{Cl}^- + \text{H}_2\text{O}$   
 reaction 2  $2\text{ClO}^- + 2\text{OH}^- \rightarrow \text{ClO}_3^- + \text{Cl}^- + \text{H}_2\text{O}$   
*water doesn't react*
- D** reaction 1  $\text{Cl}_2 + 2\text{OH}^- \rightarrow \text{ClO}^- + \text{Cl}^- + \text{H}_2\text{O}$   
 reaction 2  $3\text{ClO}^- \rightarrow \text{ClO}_3^- + 2\text{Cl}^-$



- 18 Which statement explains the observation that magnesium hydroxide dissolves in aqueous ammonium chloride, but not in aqueous sodium chloride?

- ~~A~~ The ionic radius of the  $\text{NH}_4^+$  ion is similar to that of  $\text{Mg}^{2+}$  but not that of  $\text{Na}^+$ .
- ~~B~~  $\text{NH}_4\text{Cl}$  dissociates less fully than  $\text{NaCl}$ . *contradicts with the above statement*
- ~~C~~ The  $\text{Na}^+$  and  $\text{Mg}^{2+}$  ions have the same number of electrons. *If they have same no. of outer electrons they should be able to react the same way.*
- D** The  $\text{NH}_4^+$  ion can donate a proton.



- 19 Transition elements and their compounds are widely used as catalysts.

What is the identity and what is the oxidation number of the element present in the catalyst used in the Contact process? *Vanadium(V) oxide*

	element	oxidation number
A	iron	0
B	iron	+3
C	vanadium	0
<b>D</b>	vanadium	+5

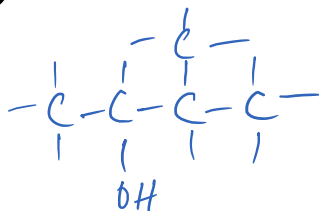
- 20 What is true of **every** nucleophile? *donates electron pair*

- ~~A~~ It attacks a double bond. *that's electrophile's job*
- B** It donates a lone pair of electrons.
- ~~C~~ It is a single atom. *HCN is also a nucleophile*
- ~~D~~ It is negatively charged. *//*

- 21 X has the molecular formula  $C_5H_{12}O$ . X has a branched carbon skeleton and a secondary alcohol functional group.

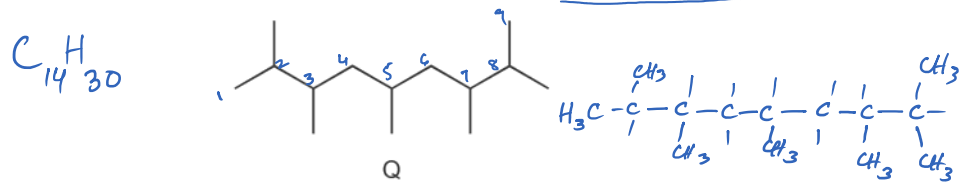
How many **structural** isomers fit this description of X?

- A** 1                      B 2                      C 3                      D 4

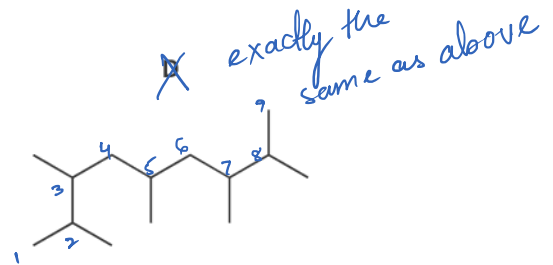
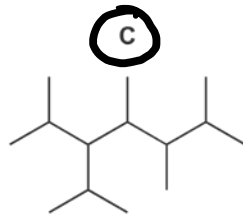
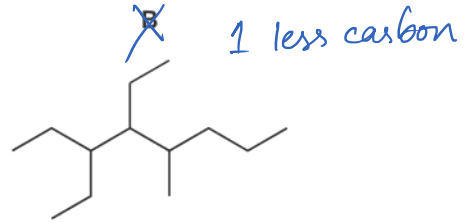
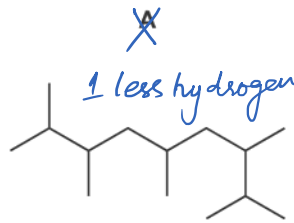




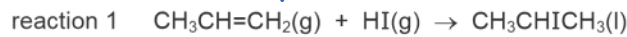
- 22 A new jet fuel has been produced that is a mixture of different structural isomers of compound Q.



Which skeletal formula represents a **structural isomer** of Q?

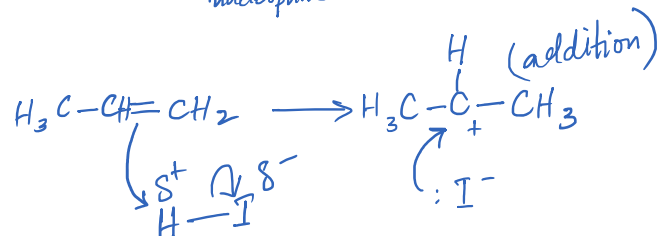
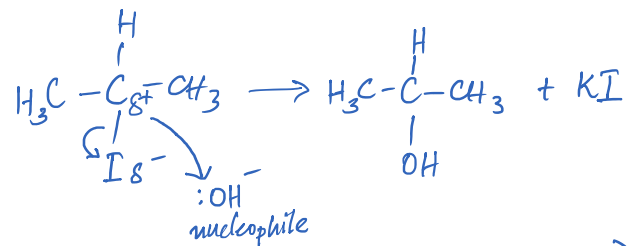


- 23 The conversion of propene to propan-2-ol can be carried out in two stages represented by the equations shown.



How can these two reactions be described?

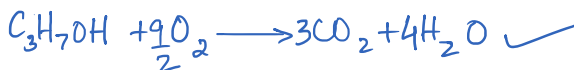
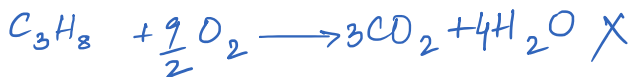
	reaction 1	reaction 2
A	addition	elimination
<b>B</b>	addition ✓	substitution ✓
C	elimination	substitution ✓
D	substitution	elimination



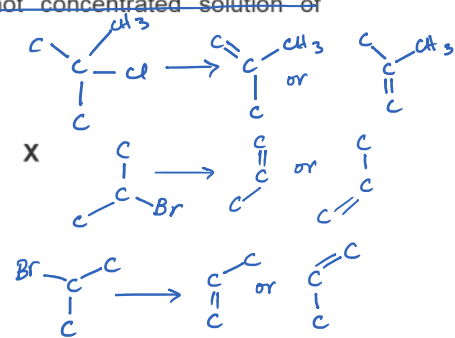
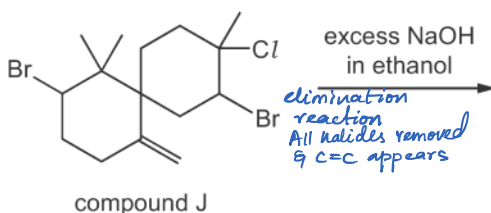
- 24 An organic molecule W contains 3 carbon atoms. It requires 4.5 molecules of oxygen for complete combustion.

What could W be?

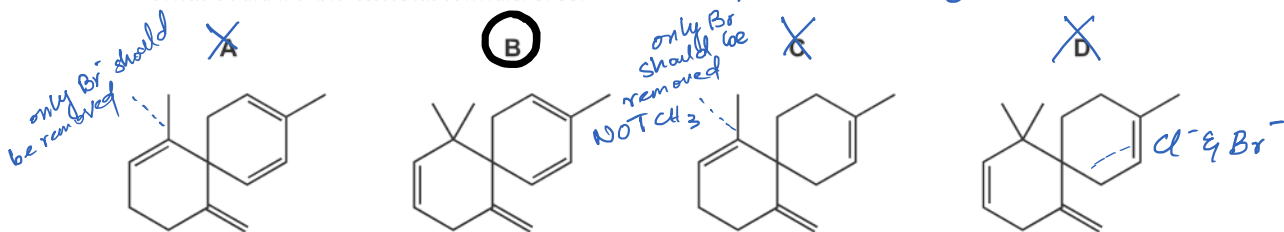
- A propane
- B propanoic acid
- C propanone
- D** propan-1-ol



- 25 Compound J,  $C_{15}H_{23}Br_2Cl$ , is reacted with an excess of a hot concentrated solution of sodium hydroxide in ethanol. One of the products is X.



What could be the skeletal formula of X?



- 26 Sodium reacts with 1 mol of compound Y to produce 1 mol of  $H_2(g)$ .  
 $Na + \text{alcohol} \rightarrow H_2(g)$   
 $Na + \text{acid} \rightarrow H_2(g)$

Which compound could Y be?

- A  $CH_3CH_2CH_2CH_2OH$   $C_4H_9OH + Na \rightarrow C_4H_9ONa + \frac{1}{2} H_2 \quad \times$
- B  $(CH_3)_3COH$   $(CH_3)_3COH + Na \rightarrow (CH_3)_3CONa + \frac{1}{2} H_2$
- C  $CH_3CH_2CH_2CO_2H$   $C_3H_7COOH + Na \rightarrow C_3H_7COONa + \frac{1}{2} H_2$
- D**  $CH_3CH(OH)CO_2H$   $CH_3CH(OH)CO_2H + Na \rightarrow CH_3CH(ONa)CO_2Na + H_2$

- 27 Which compound shows optical isomerism and gives a positive test with alkaline aqueous iodine?

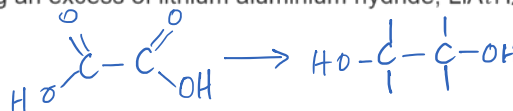
- A**  $CH_3COCH(OH)CH_3$
- B  $CH_3COCH_2CH_2OH \quad \times$
- C  $HOCH_2CH(CH_3)CHO \quad \times$
- D  $(CH_3)_2C(OH)CHO \quad \times$

ethanal / secondary alcohol / ketone

acid reduces to primary <sup>11</sup> alcohol

- 28 Ethanedioic acid, HO<sub>2</sub>CCO<sub>2</sub>H, is reduced using an excess of lithium aluminium hydride, LiAlH<sub>4</sub>.

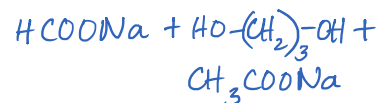
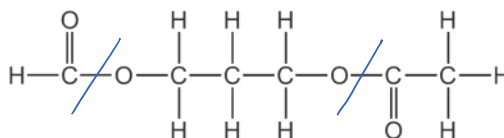
What is the organic product of the reaction?



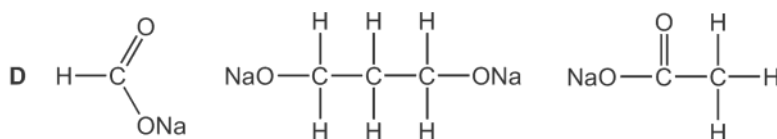
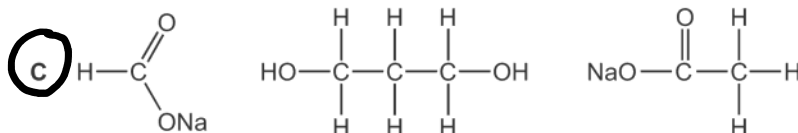
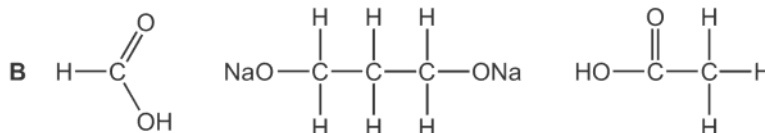
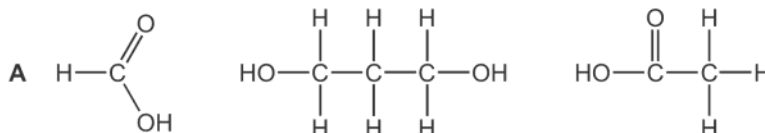
- A ethanol  
**B ethane-1,2-diol**  
 C ethanedial, OHCCHO  
 D methane

- 29 The diester shown can be hydrolysed by heating with an excess of aqueous sodium hydroxide.

Instead of acid, salt forms along with alcohol.  
 it reacts with acid group only.

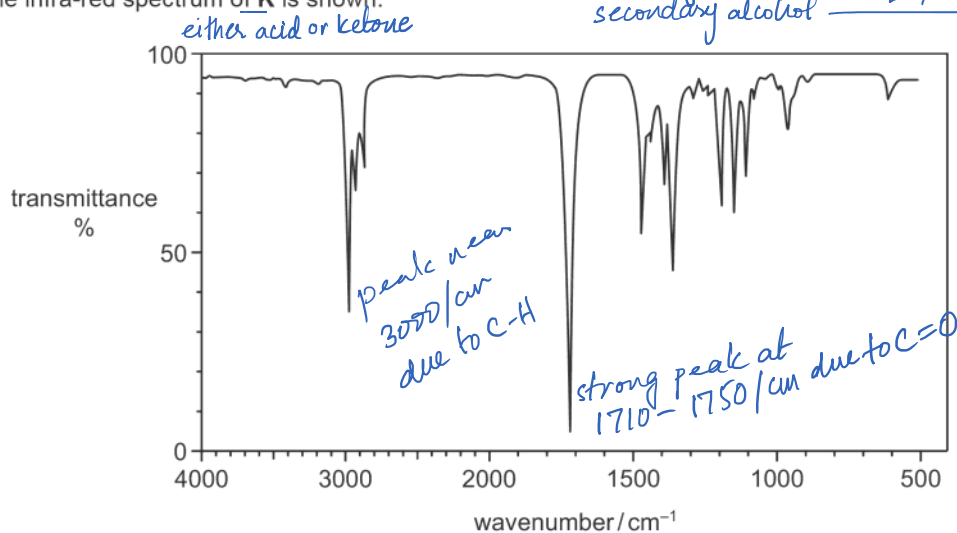


What would the products of this reaction be?



- 30 **J** is a branched-chain alcohol,  $C_5H_{12}O$ . **J** is heated under reflux with an excess of  $Cr_2O_7^{2-}/H^+$  until no further reaction occurs. An organic compound **K** is formed in good yield.

The infra-red spectrum of **K** is shown.



What are the structures of the branched-chain alcohol **J** and compound **K**?

	<b>J</b>	<b>K</b>
<del>A</del>	$CH_3CH(CH_3)CH_2CH_2OH$	$CH_3CH(CH_3)CH_2CHO$ X
<del>B</del>	$CH_3CH_2CH(OH)CH_2CH_3$ <i>not branched chain</i>	$CH_3CH_2COCH_2CH_3$
<b>C</b>	$CH_3CH(CH_3)CH(OH)CH_3$	$CH_3CH(CH_3)COCH_3$
<del>D</del>	$CH_3CH(CH_3)CH_2CH_2OH$	$CH_3CH(CH_3)CH_2COOH$ <i>no broad peak at 2500-3000/cm</i>

*BC2 it says that heating continues until no further reaction takes place*

## Section B

For each of the questions in this section, one or more of the three numbered statements 1 to 3 may be correct.

Decide whether each of the statements is or is not correct (you may find it helpful to put a tick against the statements that you consider to be correct).

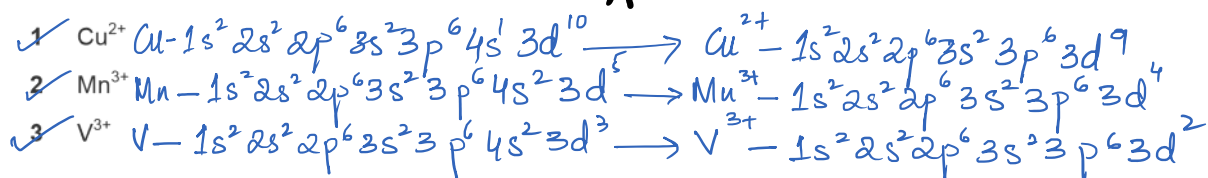
The responses **A** to **D** should be selected on the basis of

A	B	C	D
1, 2 and 3 are correct	1 and 2 only are correct	2 and 3 only are correct	1 only is correct

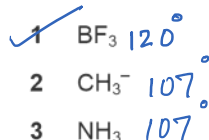
No other combination of statements is used as a correct response.

Use of the Data Booklet may be appropriate for some questions.

31 Which ions contain one or more unpaired electrons? **A**



32 Which molecules and ions have a bond angle of  $120^\circ$ ? **D**



33 Which statements are correct for **all** exothermic reactions? **B**

- 1  $\Delta H$  for the reaction is negative.  
 2 On a reaction pathway diagram the products are shown lower than the reactants.  
 3 The reaction will occur without heating. *like enthalpy of neutralisation*

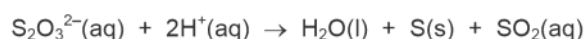


The responses **A** to **D** should be selected on the basis of

A	B	C	D
1, 2 and 3 are correct	1 and 2 only are correct	2 and 3 only are correct	1 only is correct

No other combination of statements is used as a correct response.

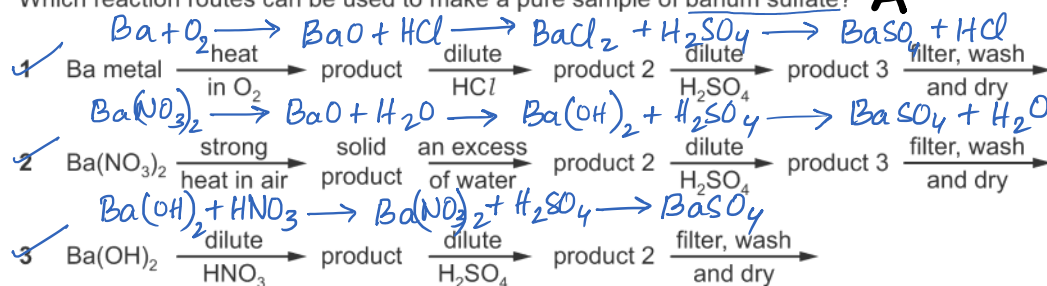
- 34 The factors affecting the rate of reaction between aqueous sodium thiosulfate and hydrochloric acid can be investigated. The ionic equation for the reaction is shown.



Which of the following can be used to investigate the rate of this reaction? **C**

- 1 change of mass
- 2 change of appearance caused by formation of a precipitate *the faster it appears that ppt is forming, the higher the rate is*
- 3 change of electrical conductivity *ions are involved in this reaction*

- 35 Which reaction routes can be used to make a pure sample of <sup>insoluble</sup> barium sulfate? **A**

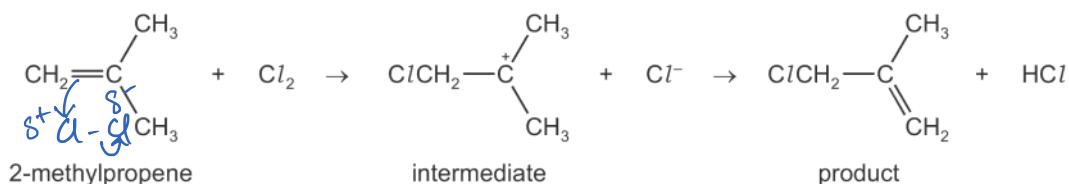


- 36 Which properties increase from magnesium to barium? **A**

- 1 ionic radius of the cation  $\text{M}^{2+}$
- 2 screening of outermost electrons by inner shells
- 3 solubility of the hydroxides,  $\text{M}(\text{OH})_2$ , in water

37 2-methylpropene can react in more than one way with chlorine.

One of the reactions follows the pathway shown.



Which statements about this mechanism are correct? **B**

- ✓ 1 The intermediate has all carbon atoms in the same plane.
- ✓ 2 There is an electrophilic attack on the double bond.
- ✗ 3 It is a free radical mechanism. *alkane + halogen → alkylhalide + ...*

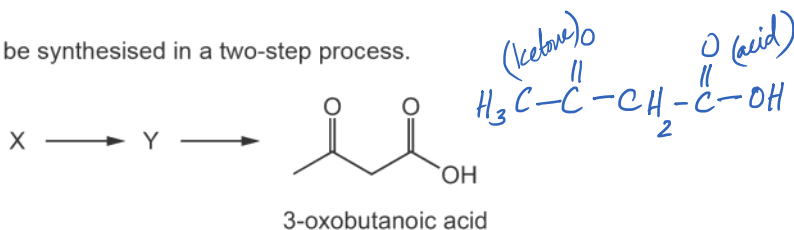
38 The halogenoalkanes listed below all react with NaOH(aq).

Which reactions proceed mainly by an S<sub>N</sub>1 mechanism? **C**

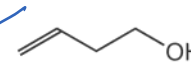
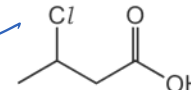
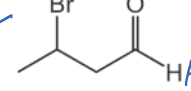
- ✗ 1 1-iodopropane
- ✓ 2 2-iodo-2-methylpropane
- ✓ 3 2-bromo-2-methylbutane

*only tertiary halogenoalkanes*

39 3-oxobutanoic acid can be synthesised in a two-step process.



What could be the structure of X? **A**

- ✓ 1   $\text{H}_2\text{C}=\text{CH}-\text{CH}_2-\text{CH}_2-\text{OH} \xrightarrow{+\text{H}_2\text{O}} \text{H}_3\text{C}-\text{CH}(\text{OH})-\text{CH}_2-\text{CH}_2-\text{OH} \xrightarrow{+\text{K}_2\text{Cr}_2\text{O}_7} \text{H}_3\text{C}-\overset{\text{O}}{\parallel}{\text{C}}-\text{CH}_2-\text{CO}_2\text{H}$
- ✓ 2   $\text{H}_3\text{C}-\text{CHCl}-\text{CH}_2-\text{CO}_2\text{H} \xrightarrow{+\text{NaOH}} \text{H}_3\text{C}-\text{CH}(\text{OH})-\text{CH}_2-\text{CO}_2\text{H} \xrightarrow{+\text{K}_2\text{Cr}_2\text{O}_7} \text{H}_3\text{C}-\text{CO}-\text{CH}_2-\text{CO}_2\text{H}$
- ✓ 3   $\text{H}_3\text{C}-\text{CHBr}-\text{CH}_2-\text{CHO} \xrightarrow{+\text{NaOH}} \text{H}_3\text{C}-\text{CH}(\text{OH})-\text{CH}_2-\text{CHO} \xrightarrow{+\text{K}_2\text{Cr}_2\text{O}_7} \text{H}_3\text{C}-\text{CO}-\text{CH}_2-\text{CO}_2\text{H}$

The responses **A** to **D** should be selected on the basis of

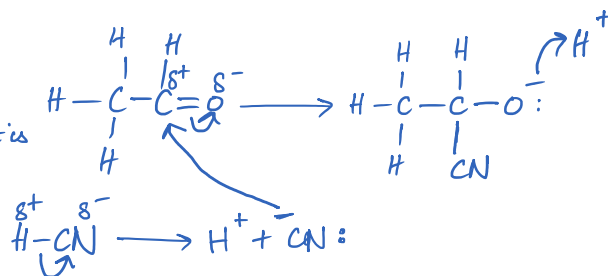
A	B	C	D
1, 2 and 3 are correct	1 and 2 only are correct	2 and 3 only are correct	1 only is correct

No other combination of statements is used as a correct response.

40 Ethanal reacts with HCN in the presence of KCN. *to form*  $\text{H}_3\text{C}-\overset{\text{H}}{\underset{\text{CN}}{\text{C}}}-\text{OH}$

Which changes in bonding occur during this reaction? **B**

- 1 A carbon-carbon bond is formed. *C-CN*
- 2 A carbon-hydrogen bond is broken. *H-CN*
- 3 A carbon-nitrogen bond is broken. *:CN<sup>-</sup> remains as it is*



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